

tices and 24 edges to the decahedron, but as basal faces are shared there are 12 vertices and 20 edges per unit of the column. There are then four hydrogen atoms excess to each unit, one *per* three water molecules, and these form hydrogen bonds to the phenolic oxygen atoms of the nickel complex. The column of water molecules, infinite in one-dimension, represents an extremely compact assembly of water molecules provided that hydrogen acceptors of appropriate number and dimension are available, but it could not be extended to a three-dimensional structure in their absence. The structure is then in no sense a clathrate or potential clathrate, but is an interesting intermediate between these compounds and more conventional hydrates in which the arrangement of water molecules, *per se*, is less extensive.

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## The Crystal and Molecular Structure of 2-Dimethylaminoethyl Selenolbenzoate Hydrochloride

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The crystal structure of the local anesthetic 2-dimethylaminoethyl selenolbenzoate hydrochloride has been determined. The crystals are monoclinic,  $P2_1/c$  with  $a=12.287$  (3),  $b=6.945$  (1),  $c=15.523$  (3) Å and  $\beta=94.57$  (2)°. Least-squares refinement on 2411 observed reflections resulted in an  $R$  value of 0.051. The  $\overset{+}{N}-C-C-Se$  group adopts the *trans* conformation with a torsion angle about the C-C bond of 174.4°. The two Se-C bond distances are both 1.945 Å with estimated standard deviations of 0.005 Å. The C-Se-C angle is 96.4 (2)°. Cell constants for four related compounds are also given.

### Introduction

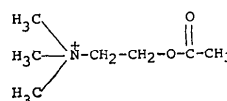
The conformations of biologically active molecules related to acetylcholine (I) have been extensively studied (Chothia, 1970; Chothia & Pauling, 1970; Shefter & Mautner, 1969). These molecules all incorporate the  $\overset{+}{N}-C-C-R_1$  group where  $R_1$  may be oxygen, sulfur or selenium. While these short chains are formally flexible, the  $\overset{+}{N}-C-C-O$  group usually adopts a *gauche* conformation (Pauling, 1968; Sundaralingam, 1968; Chothia & Pauling, 1970). Other studies show that the  $\overset{+}{N}-C-C-S$  and  $\overset{+}{N}-C-C-Se$  groups take up the *trans* conformation (Shefter & Mautner, 1969).

In this work five members of a series of local anesthetics related to the tertiary analog of benzoylcholine, 2-dimethylaminoethyl benzoate (II, with  $R_1=O$ ,  $R_2=O$ ) were investigated. These compounds are obtained by systematically replacing the ether oxygen in

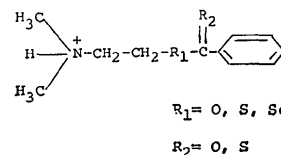
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the parent molecule (position  $R_1$ ) with sulfur or selenium and the carboxyl oxygen (position  $R_2$ ) with sulfur. Only the selenolester 2-dimethylaminoethyl selenolbenzoate hydrochloride ( $R_1=Se$ ,  $R_2=O$ ) was selected for complete structure determination.



(I)

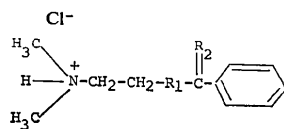


(II)

$R_1 = O, S, Se$   
 $R_2 = O, S$

Crystals of the parent compound ( $R_1=O$ ,  $R_2=O$ ) proved highly deliquescent and were not investigated. The structure of the closely related 2-diethylaminoethyl *p*-aminobenzoate HCl (procaine HCl) was determined instead and a discussion of these results will appear subsequently (Dexter, 1972).

Table 1. Cell Constants for


 $R_1 = \text{O, S, Se}$ 
 $R_2 = \text{O, S}$ 
Space Group  $P2_1/c$ 

	$R_1 = \text{S, R}_2 = \text{S}$	$R_1 = \text{Se, R}_2 = \text{S}$	$R_1 = \text{Se, R}_2 = \text{O}$	$R_1 = \text{S, R}_2 = \text{O}$	$R_1 = \text{O, R}_2 = \text{S}$
$a$ (Å)	11.81 (1)	11.95 (1)	12.26 (1)	12.15 (1)	9.73 (1)
$b$ (Å)	6.90 (1)	6.85 (1)	6.98 (1)	6.93 (1)	9.22 (1)
$c$ (Å)	16.68 (1)	16.66 (1)	15.43 (2)	15.34 (1)	12.74 (1)
$\beta$ (°)	92.8 (2)	92.9 (2)	94.8 (2)	94.4 (2)	115.2 (2)

### Experimental

Shown in Table 1 are the approximate cell dimensions of the five compounds as measured from Weissenberg and precession photographs. In all cases systematic absences  $h0l$  for  $l$  odd and  $0k0$  for  $k$  odd uniquely determine the space group as  $P2_1/c$ .

A crystal of the selenolester, approximately  $0.18 \times 0.20 \times 0.14$  mm, was cut from a large, colorless plate and mounted with  $b$  along the  $\varphi$  axis of a General Electric computer-controlled diffractometer. The  $2\theta$  values (Cu  $K\alpha_1$ ,  $\lambda = 1.54050$  Å) of 15 carefully centered high angle reflections were used in a least-squares refinement of the lattice constants. (Crystal data are presented in Table 2.) The intensities of the 2810 independent reflections with  $2\theta < 155^\circ$  (Cu  $K\alpha$ ,  $\lambda = 1.5418$  Å) were measured using  $\theta$ - $2\theta$  scans. The 399 reflections for which  $I_{\text{obs}} < 2\sigma(I_{\text{obs}})$  were all treated as unobserved and assigned an intensity equal to  $2\sigma(I_{\text{obs}})$ . The intensities were corrected for Lorentz and polarization factors but not for absorption.

During data collection the intensities of the three standard reflections decreased linearly at rates which seemed to be a function of  $2\theta$ . The crystal also became yellow-brown in color (possibly due to liberation of

free selenium). Although the losses in intensity (9 to 33%) were significant, the solution of the structure proceeded uneventfully and exploratory refinement cycles led to a chemically reasonable structure.

Table 2. Crystal data for 2-dimethylaminoethyl selenolbenzoate hydrochloride

Formula	$(\text{CH}_3)_2\text{N}(\text{CH}_2)_2\text{Se COC}_6\text{H}_5 \cdot \text{HCl}$
M. W.	292.7
Space group	$P2_1/c$
$a$	12.387 (3) Å
$b$	6.945 (1)
$c$	15.523 (3)
$\beta$	94.57 (2)°
Volume	1331.2 Å <sup>3</sup>
$\rho$ (X-ray)	1.460 g.cm <sup>-3</sup>
$\rho$ (floatation)	1.47 g.cm <sup>-3</sup>
$Z$	4
$\mu$	59.3 cm <sup>-1</sup>

The estimated standard deviations of the lattice constants are three times those obtained from the least-squares calculation.

### Structure determination and refinement

The selenium atom and chloride ion were readily located from an  $E^2 - 1$  vector map. Coordinates and

Table 3. Final positional and anisotropic thermal parameters for the non-hydrogen atoms

The positional parameters for  $\text{Cl}^-$  and Se are  $\times 10^5$ . All others are  $\times 10^4$ . The estimated standard deviations are those obtained from the least-squares and refer to the last decimal place given. The form of the anisotropic temperature factor is  $\exp[-\frac{1}{4}(B_{11}h^2a^*2 + B_{22}k^2b^*2 + B_{33}l^2c^*2 + 2B_{12}hka^*b^* + 2B_{13}hla^*c^* + 2B_{23}klb^*c^*)]$ .

	$x$	$y$	$z$	$B_{11}$	$B_{22}$	$B_{33}$	$B_{12}$	$B_{13}$	$B_{23}$
C(1)	3598 (4)	1948 (6)	7904 (3)	4.9 (2)	4.4 (2)	3.1 (1)	-0.2 (1)	-0.7 (2)	0.1 (1)
C(2)	5371 (3)	2051 (7)	8721 (3)	3.8 (1)	4.6 (2)	5.0 (2)	0.3 (1)	0.7 (2)	0.4 (1)
C(3)	3744 (2)	1892 (4)	9518 (2)	3.7 (1)	2.7 (1)	3.0 (1)	0.0 (1)	-0.2 (1)	0.6 (1)
C(4)	2635 (3)	2699 (6)	9633 (2)	3.4 (1)	4.2 (2)	3.7 (1)	0.4 (1)	0.2 (1)	1.0 (1)
C(5)	8458 (2)	4451 (5)	4650 (2)	3.4 (1)	4.5 (2)	2.9 (1)	0.3 (1)	0.2 (1)	0.2 (1)
C(6)	8747 (2)	3128 (5)	3950 (2)	3.0 (1)	3.9 (1)	2.9 (1)	-0.3 (1)	0.2 (1)	0.1 (1)
C(7)	8210 (3)	3159 (5)	3128 (2)	3.0 (1)	4.4 (2)	3.1 (1)	-0.0 (1)	-0.1 (1)	0.0 (1)
C(8)	8510 (3)	1892 (5)	2500 (2)	3.7 (1)	5.1 (2)	3.4 (1)	-0.8 (1)	-0.1 (1)	-0.7 (1)
C(9)	9353 (3)	0642 (6)	2680 (2)	4.6 (2)	4.5 (2)	4.5 (2)	-0.1 (1)	0.5 (2)	-1.0 (1)
C(10)	9894 (3)	0580 (6)	3504 (3)	4.2 (1)	4.4 (2)	5.2 (2)	0.8 (1)	-0.0 (2)	-0.3 (1)
C(11)	9587 (3)	1804 (5)	4133 (2)	4.1 (1)	4.3 (2)	3.7 (2)	0.2 (1)	-0.6 (1)	0.2 (1)
N	4207 (2)	2611 (4)	8711 (2)	3.5 (1)	2.4 (1)	2.9 (1)	-0.1 (1)	-0.1 (1)	0.4 (1)
O	8635 (3)	4106 (5)	5406 (2)	7.3 (2)	6.2 (2)	2.6 (2)	1.6 (1)	0.1 (1)	0.1 (1)
$\text{Cl}^-$	58321 (8)	19387 (1)	62870 (6)	5.35 (4)	2.22 (3)	4.45 (4)	-0.08 (2)	-0.62 (3)	-0.02 (2)
Se	21848 (3)	18784 (6)	07473 (2)	4.98 (2)	4.76 (2)	3.64 (2)	-0.51 (1)	0.47 (2)	-0.25 (1)

isotropic thermal parameters for the heavy atoms and the scale factor were adjusted with one cycle of full-matrix least squares (with unit weights) to give an  $R$  value of 0.35. The subsequent electron density map revealed the other 13 non-hydrogen atoms. After several least-squares cycles with individual isotropic temperature factors  $R$  decreased to 0.12. At this point the hydrogen atoms were located from a difference map and included in the refinement with fixed isotropic temperature factors. After allowing the heavy atoms to refine anisotropically  $R$  was 0.075. Weights were then assigned from a plot of  $|\Delta F|$  vs.  $|F_o|$ . Several intense low angle reflections appeared to be affected by extinction and were excluded from the remainder of the refinement. Using anisotropic thermal parameters for all non-hydrogen atoms and isotropic thermal parameters for hydrogen atoms, the refinement converged with  $R=0.051$ . In the last cycle for the 200 parameters refined, the average change was  $0.04\sigma$  and the maximum change was  $0.4\sigma$ . The final difference map showed a peak of about 0.5 electron at the position of the chloride ion and a pattern of peaks and

holes about the selenium atom position with heights of about  $\pm 1.0$  electron. While this pattern might suggest incomplete refinement of the anisotropic thermal parameters, it also could be a manifestation of the crystal decay noted during the data collection.

Scattering factors for carbon, nitrogen and oxygen were taken from *International Tables for X-ray Crystallography* (1962), for selenium and chlorine ( $\text{Cl}^-$ ) from Cromer & Waber (1964) and for hydrogen from Stewart, Davidson & Simpson (1965). The bulk of the computations was carried out with the *X-ray 67 System* programs (Stewart, 1967). The mean plane program used was that of Ahmed, Hall, Pippy & Saunderson (1966).

Table 4. Final coordinates and isotropic thermal parameters for the hydrogen atoms

The positional parameters are all  $\times 10^3$ .

	$x$	$y$	$z$	$B$
H(1)	356 (5)	050 (9)	792 (4)	8 (1)
H(2)	379 (4)	239 (8)	736 (4)	7 (1)
H(3)	309 (3)	217 (6)	787 (3)	3 (1)
H(4)	545 (4)	070 (9)	866 (3)	6 (1)
H(5)	574 (4)	262 (7)	919 (3)	5 (1)
H(6)	560 (4)	245 (8)	830 (3)	5 (1)
H(7)	377 (3)	061 (6)	957 (2)	3 (1)
H(8)	420 (2)	217 (5)	995 (2)	3 (1)
H(9)	205 (5)	242 (9)	926 (4)	7 (1)
H(10)	269 (3)	405 (7)	962 (3)	4 (1)
H(11)	761 (3)	398 (6)	303 (3)	4 (1)
H(12)	825 (4)	194 (7)	198 (3)	6 (1)
H(13)	956 (4)	-018 (7)	225 (3)	6 (1)
H(14)	046 (4)	-018 (8)	369 (3)	7 (1)
H(15)	002 (4)	188 (6)	465 (3)	6 (1)
H(16)	418 (3)	410 (6)	871 (2)	3 (1)

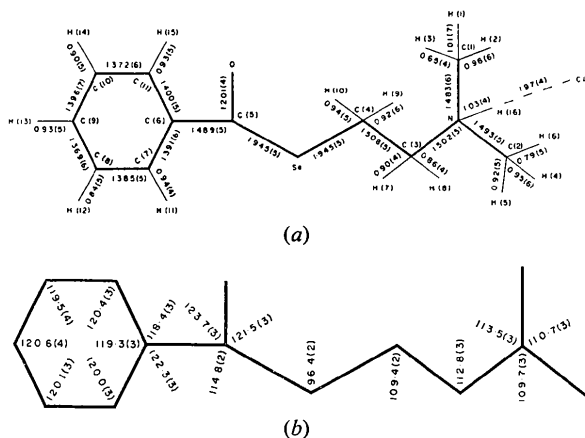


Fig. 1. (a) Bond distances in 2-dimethylaminoethyl selenolbenzoate hydrochloride. (b) Bond angles involving non-hydrogen atoms. (Angles involving H atoms are within expected limits.)

Table 5. Torsion angles

	Angle ( $\tau$ )
C(11)-C(6)-C(5)-Se	153.2°
C(6)-C(5)-Se-C(4)	169.4
O-C(5)-Se-C(4)	166.9
C(5)-Se-C(4)-C(3)	87.4
Se-C(4)-C(3)-N	174.4
C(4)-C(3)-N-C(2)	169.7
C(4)-C(3)-N-C(1)	65.9

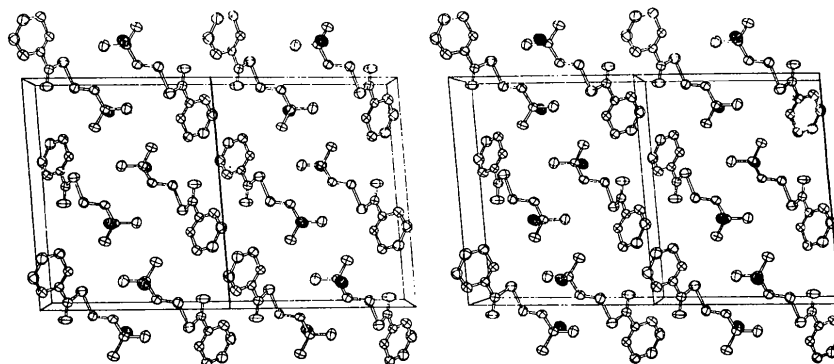


Fig. 2. Stereo diagrams (Johnson, 1965) showing molecular packing along  $b$ .  $a \rightarrow$ ,  $c \uparrow$ . Hydrogen atoms have been omitted. Chloride ions have been shaded.

Results and discussion

Final coordinates and thermal parameters for non-hydrogen atoms are given in Table 3 and for hydrogen

atoms in Table 4. Fig. 1 shows bond distances, bond angles and the numbering system. Some torsion angles along the side chain are listed in Table 5. Table 6 is the list of observed and calculated structure factors.

Table 6. Observed and calculated structure factors

The columns are *l*, 10*F*<sub>o</sub> and 10*F*<sub>c</sub>. Unobserved reflections are marked with an asterisk. Reflections marked with an *E* were excluded from the final refinement cycles.

Table with multiple columns containing numerical data for structure factors, including observed (F\_o) and calculated (F\_c) values for various reflections. The table is organized into several sections based on reflection indices (h, k, l).



significantly less stable than the *trans*. In the present structure, the observed torsion angle probably arises to accommodate the formation of different environments for the alkyl and aromatic parts of the molecule. A packing diagram (Fig. 2) shows that the protonated nitrogen atoms are hydrogen bonded to chloride ions (N-Cl, 3.07 Å) so that layers of charged groups alternate with layers of interlocking phenyl rings.

The smoothness of the refinement and the reasonableness of the final structure prompted an attempt to assess the effect of the crystal decay on the results.

The  $hk0$  reflections were measured twice: at the beginning of the experiment as part of the  $hkl$  reflections ( $I_0$ ) and later as part of the  $hkl$  reflections ( $I_{1/2}$ ). Fig. 3 is a plot of  $\ln(I_{1/2}/I_0)$  versus  $\sin^2\theta/\lambda^2$ . Assuming that the decrease in intensity can be described in terms of an apparent  $B$  parameter, by the end of the data collection the total effect of the decomposition was an increase in the overall temperature factor by about  $1.5 \text{ \AA}^2$ .

The effect of a larger temperature factor is to increase the uncertainty in atomic positions, not to displace them (Cruickshank, 1949). It seems reasonable, therefore, to accept as valid the positional parameters ob-

tained from the data affected by crystal decay – with the usual warning that the standard deviations obtained from the least-squares probably will be too small.

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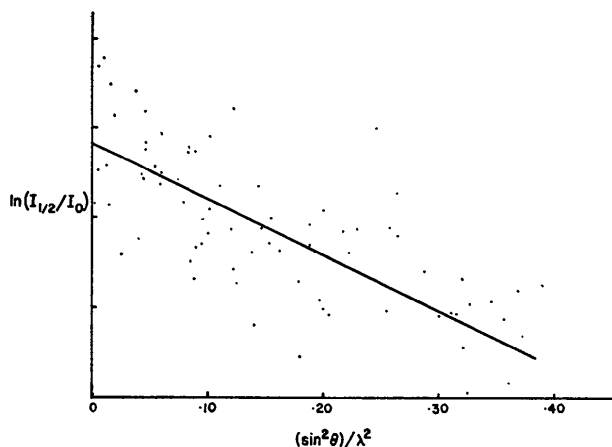


Fig. 3. Plot of  $\ln(I_{1/2}/I_0)$  versus  $(\sin^2\theta)/\lambda^2$  showing loss of intensity at midpoint of data collection for 75  $hk0$  reflections. The slope of the least-squares straight line corresponds to a  $B$  parameter of  $0.8 (1) \text{ \AA}^2$ .